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First quickly scan the worksheet and circle any metals (as a symbol or as a name) that is a transition

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metal. Then either give the same or chemical formula as necessary for each problem below. Remember that transition metals can have multiple oxidation states, so you are required

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## ~~Answer . . .~~ Worksheet Answers

1. Write the chemical formulas for the following binary ionic compounds:  
barium oxide  $\text{BaO}$  lithium sulfide  $\text{Li}_2\text{S}$  magnesium bromide  $\text{MgBr}_2$  strontium iodide  $\text{SrI}_2$  calcium sulfide

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CaS hydrogen bromide HBr  
lithium oxide Li<sub>2</sub>O potassium  
chloride KCℓ cadmium  
fluoride CdF<sub>2</sub> silver sulfide  
Ag<sub>2</sub>S potassium phosphide K

~~Nomenclature #1: Binary  
Ionic Compounds~~

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A B; lithium fluoride: LiF:  
lithium chloride: LiCl:

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lithium bromide: LiBr:  
lithium iodide: LiI: ...

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Compounds contain Transition  
metals (including the Group



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III, IV, V, VI metals,  
except for Al) with non-  
metal ions. Show the correct  
name for the following  
compounds. Give correct  
formulas for these Type II  
binary compounds

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more than one atom in the  
mixture.

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Name the following ionic compounds. Be sure to check for Roman Numerals!  $\text{Na}_3\text{P}$ .  
 $\text{MgBr}_2$ .  $\text{Ag}_2\text{Se}$ .  $\text{AlCl}_3$ .  $\text{Ca}_3\text{P}_2$ .  
 $\text{Ba}_3\text{N}_2$ .  $\text{Li}_2\text{O}$ .  $\text{ZnF}_2$ .  $\text{K}_2\text{S}$ .  $\text{SrI}_2$   
 $(\text{NH}_4)_2\text{S}$   $\text{Ca}_3(\text{PO}_3)_2$   $\text{Li}_2\text{CO}_3$   
 $\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3$   $\text{KCN}$   $\text{AgNO}_3$



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Sb(NO<sub>3</sub>)<sub>5</sub> NiS Cu(NO<sub>3</sub>)<sub>2</sub> FeCl<sub>2</sub>  
Au(CN)<sub>3</sub> Pb(SO<sub>4</sub>)<sub>2</sub> Write the  
correct formulas for each of  
the ionic compounds.

~~WRITING AND NAMING BINARY  
COMPOUNDS WORKSHEET~~

Portage HS Chem Combined

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always have 2 types of

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atoms. 3. When naming an ionic compound, you should name the metal first and the nonmetal second (metal or nonmetal). 4. Circle one: when naming a binary ionic compound, the ending of the first element does/does not

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change. 5. When naming a binary ionic compound the nonmetal always has an -ide ending.

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Nomenclature Worksheet 5:

Ionic Compounds Summary Name  
the following compounds:

Give the formula for each

compound: 23. sodium fluoride

24. potassium sulfide ca COQ

25. calcium carbonate 26.

magnesium hydroxide (C) 27

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zinc nitrate 28. silver  
acetate 2. 3. 4. 5. 6. 8. 9.  
CaF<sub>2</sub> BaS cuS<sub>04</sub> Fe<sub>203</sub> HgC<sub>12</sub>  
AgN<sub>03</sub> MgC<sub>03</sub> KC<sub>2H302</sub>  
29. copper (II) oxide

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the Group III, IV, V, VI  
metals, except for Al) with  
non-metal ions. Show the  
correct name for the  
following compounds.

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