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This reaction follows the equation: $N_2 + 3H_2 \rightarrow 2NH_3$. $\Delta H_f^\circ = -92 \text{ kJ mol}^{-1}$, $\Delta S_f^\circ = -192 \text{ J mol}^{-1} \text{ K}^{-1}$. To find the Gibbs free energy of formation at room temperature, recall that $\Delta G_f^\circ = \Delta H_f^\circ - T\Delta S_f^\circ = -92 \text{ kJ mol}^{-1} + (298 \text{ K})(0.192 \text{ kJ mol}^{-1} \text{ K}^{-1}) = -35 \text{ kJ mol}^{-1}$. Alternatively, one can find the temperature for which $\Delta G_f^\circ = 0$, $T = \Delta H_f^\circ / \Delta S_f^\circ$.

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One of the most difficult problems in chemical reaction engineering is to navigate in a reaction network efficiently in order to optimize the production of the desired target component (Problem 4). In this field again catalysis is a main tool.

1 Basic Problems of Chemical Reaction Engineering and ...

example problems before going on to other solutions. P2-5. This is a reasonably challenging problem that reinforces Levenspiel's plots. P2-6. Straight forward problem alternative to problems 7, 8, and 11. P2-7. To be used in those courses emphasizing bio reaction engineering. P2-8. The answer gives ridiculously large reactor volume.

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