

Chemistry B Moles Packet Answer Key

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7.B Mole and Mass Avogadro's Number, The Mole, Grams, Atoms, Molar Mass Calculations - Introduction Converting Between Grams and Moles Mole Ratio Practice Problems **Concept of Mole - Part 1 | Atoms and Molecules | Don't Memorise** Empirical Formula Molecular Formula Determination From Percent Composition **Molality Practice Problems - Molarity, Mass Percent, and Density of Solution Examples** ~~Introduction to Moles~~ Stoichiometry Basic Introduction, Mole to Mole, Grams to Grams, Mole Ratio Practice Problems *Mole Conversions Made Easy: How to Convert Between Grams and Moles* How To Convert Grams To Moles - VERY EASY! Counting Atoms: Intro to Moles Part 2 ~~Converting Grams to Moles Using Molar Mass | How to Pass Chemistry~~ **The Mole: Avogadro's Number and Stoichiometry** Introduction to Combustion Analysis, Empirical Formula Molecular Formula Problems Dalton's Law of Partial Pressure Problems Examples - Chemistry Stoichiometry ~~Mole to Mole Conversions~~ ~~Molar Ratio Practice Problems~~ Step by Step Stoichiometry Practice Problems | ~~How to Pass Chemistry~~ ~~Balancing Chemical Equations Practice Problems~~ ~~Concept~~

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~~of Mole | Avogadro's Number | Atoms and Molecules | Don't Memorise~~ *Chemistry B Moles Packet Answer*

CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 9 CHEMISTRY WORKSHEET # 6 MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP.

CHEMISTRY B- MOLES PACKET NAME: HR: _____ PAGE 1

CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 4 CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

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A mole of carbon contains 6.02×10^{23} atoms of carbon. A mole of helium contains 6.02×10^{23} atoms of helium. A mole of sodium contains 6.02×10^{23} atoms of sodium. A mole of gold contains 6.02×10^{23} atoms of gold. However, we have learned about seven elements that exist as diatomic molecules—H₂, N₂, O₂, F₂, Cl₂, Br₂, and I₂

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CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 4 CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique. 1 mol = a molar mass of an atom/molecule (g/mol) EXAMPLE: What is the mass of 5.00 moles of water H₂O? STEP 1: 2 H = 2 x (1.01) = 2.02 STEP 2: ? grams H₂O = 5.00 moles H₂O x 18.02 g H₂O = 90.10 ...

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CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 6 CHEMISTRY WORKSHEET # 4 MOLAR MASSES AND AVOGADRO'S NUMBER Now that you know two definitions of a mole (a molar mass and an Avogadro's number of particles) you can combine these two definitions into one

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problem.EXAMPLE: How many molecules are there in 90.1 grams of water? $2 \text{ H} = 2 \times (1.01) = 2.02$ $\text{O} = 1 \times (16.00) = 16.00$ $\text{H}_2\text{O} = 18.02 \text{ g}$? molecules $\text{H}_2\text{O} = 90.1 \text{ g H}_2\text{O} \times 1 \text{ mole H}_2\text{O} \times 6.02 \times 10^{23}$ molecules $\text{H}_2\text{O} = 3.01 \times 10^{24}$ molecules $\text{H}_2 \dots$

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