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11. (a) The high/low solubility cutoff point is 0.10 mol/L at SATP. (b) The 0.10 mol/L cutoff is useful because most laboratory solutions are within an order of magnitude of this

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11(s), (sucrose) because it dissolves well, to produce a noncon-ducting, neutral solution. Chemical 2 is the ionic compound, KCl (s), because it dissolves well, to produce a conducting, neutral solution. Chemical 4 is the ionic hydroxide, Ba(OH) 2(s), because it dissolves well, to produce a conducting, basic solution.

**8.1 UNDERSTANDING ACIDS AND BASES**

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